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 We use this protocol and it's working

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Phosphate Buffered Saline (PBS)

In 1 collection

Neilier Junior¹

¹Universidade Federal de Viçosa

Neilier Junior

Universidade Federal de Viçosa, University of Manitoba

ABSTRACT

A buffer solution has the function of resisting changes in pH even when adding powerful acids or bases. However, in the physiological environment the buffered system also provides cofactors for enzymatic reactions, critical salts and even essential nutrients for cells and tissues. Therefore, when trying to reproduce biological conditions in vitro, we must make the appropriate choice of the buffer. After all, it will provide the appropriate medium in which reactions will occur.

MATERIALS

- Distilled Water
- pH Meter (sensitive)
- NaCl
- Potassium Phosphate Buffer
- Na₂HPO₄
- KH₂PO₄

SAFETY WARNINGS

! Wear personal protective equipment: gloves, lab coat and mask.

BEFORE START INSTRUCTIONS

Organize your workspace.

Make sure all solutions and equipment are available.

Phosphate Buffered Saline (PBS)

1

Note

7.4

150 mM NaCl

10 mM Potassium Phosphate buffer

Prepare 1 L PBS by dissolving 8.7 g NaCl , 1.82 g K₂HPO₄ • 3H₂O and 0.23 g KH₂PO₄ in 1 L distilled water .

1.1 A variation of PBS can also be prepared as follows:

137 millimolar (mM) NaCl

2.7 millimolar (mM) KCl

10 millimolar (mM) Na₂HPO₄

1.76 millimolar (mM) KH₂PO₄

2 Adjust the pH before use.