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# Phosphate Buffered Saline (PBS)

 In 1 collection

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Neillier Junior<sup>1</sup>

<sup>1</sup>University of Manitoba



**Neillier Junior**

University of Manitoba, Universidade Federal de Viçosa



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**Protocol status:** Working

**We use this protocol and it's working**

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## Abstract

A buffer solution has the function of resisting changes in pH even when adding powerful acids or bases. However, in the physiological environment the buffered system also provides cofactors for enzymatic reactions, critical salts and even essential nutrients for cells and tissues. Therefore, when trying to reproduce biological conditions in vitro, we must make the appropriate choice of the buffer. After all, it will provide the appropriate medium in which reactions will occur.

## Materials

- Distilled Water
- pH Meter (sensitive)
- NaCl
- Potassium Phosphate Buffer
- $\text{Na}_2\text{HPO}_4$
- $\text{KH}_2\text{PO}_4$

## Safety warnings

 Wear personal protective equipment: gloves, lab coat and mask.

## Before start

Organize your workspace.

Make sure all solutions and equipment are available.



## Phosphate Buffered Saline (PBS)






1

### Note

 7.4

150 mM NaCl

10 mM Potassium Phosphate buffer

Prepare  1 L PBS by dissolving  8.7 g NaCl ,  1.82 g  $\text{K}_2\text{HPO}_4 \cdot 3\text{H}_2\text{O}$  and  0.23 g  $\text{KH}_2\text{PO}_4$  in  1 L distilled water .

1.1 A variation of PBS can also be prepared as follows:

[M] 137 millimolar (mM) NaCl

[M] 2.7 millimolar (mM) KCl

[M] 10 millimolar (mM)  $\text{Na}_2\text{HPO}_4$

[M] 1.76 millimolar (mM)  $\text{KH}_2\text{PO}_4$

2 Adjust the pH before use.